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99	99
100	100

8. Which of the following numbers has the wrong number of significant digits?
 A. 0.78 / 208 + 183
 B. 2.08 / 0.014
 C. 6.30 - 5.9 / 2.00 + 10^2 = 3.2
9. Which of the following cannot be separated into other substances by physical processes, but can be decomposed into other substances by chemical processes, is a
 A. compound
 B. an element
 C. a mixture of elements
 D. a heterogeneous mixture
 E. a homogeneous mixture

10. Fill in the gaps in the following table:

Element	^{35}Cl	^{37}Cl
Protons	17	17
Neutrons	18	20
Electrons	17	17
Mass number	35	37
Net charge	0	0

11. Convert 8.07° to Celsius and Kelvin.
 $K = 273.15 + 8.07 = 281.22$
 $^{\circ}\text{C} = 8.07 \times \frac{5}{9} = 4.51$

12. A car is moving at 200 miles per hour. How many minutes will it take to travel 100 miles?
 $\frac{200 \text{ miles}}{1 \text{ hour}} = \frac{100 \text{ miles}}{x \text{ hours}}$
 $x = \frac{100 \text{ miles} \times 1 \text{ hour}}{200 \text{ miles}} = 0.5 \text{ hours} = 30 \text{ minutes}$

13. Chlorine has a density of 3.214 g/L. Suppose you have 500 mg of chlorine. What volume in liters will the chlorine occupy?
 $\frac{500 \text{ mg}}{1 \text{ g}} \times \frac{1 \text{ L}}{3.214 \text{ g/L}} = \frac{0.5 \text{ g}}{3.214 \text{ g/L}} = 0.156 \text{ L}$

14. The specific heat of copper is 0.385 J/g°C. How much heat energy in kilojoules is required to raise the temperature of 100 g of copper from 25°C to 75°C?
 $q = m \cdot c \cdot \Delta T = 100 \text{ g} \cdot 0.385 \text{ J/g}^{\circ}\text{C} \cdot (75 - 25)^{\circ}\text{C} = 1925 \text{ J} = 1.925 \text{ kJ}$

Covalent Compounds Worksheet

- Based on the properties of the following materials, determine whether they are made of primarily ionic compounds or covalent compounds:
 - telephone receiver: _____
 - concrete: _____
 - gasoline: _____
 - candy corn: _____
- Name the following covalent compounds:
 - SiF₄: _____
 - N₂O: _____
 - HBr: _____
 - Br₂: _____
- Write the formulas for the following covalent compounds:
 - diboron hexahydride: _____
 - nitrogen tribromide: _____
 - sulfur hexachloride: _____
 - diphosphorus pentoxide: _____
- Write the empirical formulas for the following compounds:
 - C₄H₁₀: _____
 - boron trichloride: _____
 - methane: _____
 - C₆H₁₂O₆: _____
- List three differences between ionic and covalent compounds:

GAGE Worksheet for Molecular Geometry, Polarity, Hybridization

1. Complete the following table. Use N/A where something does not apply. Please use dots for all pairs; because of the drawing application used the bonds cannot be displayed as electron pairs on this answer sheet

Compound/ion	Dot Structure	Molecular Geometry (name only, no drawings)	Polar or Non-Polar Molecule?	Type of Hybridization
NHCl ₂		trigonal pyramidal	polar, electron pair and chlorine are negative, H is positive	sp ³
MgS	[Mg ²⁺][S ²⁻]	N/A, ionic so ions are separate	N/A, ionic, not covalent molecule so no electron pair sharing	N/A
CH ₄ F		tetrahedral	polar with F as negative end of dipole	sp ³
BrCl ₃		T-shaped	polar with electron pairs as negative end of dipole	sp ³ or sp ^{3d}
SiO ₄ ²⁻		trigonal planar	N/A, ion with charge	sp ²

VSEPR Theory Predicting the 3-D Shapes of Molecules

Gather together the following equipment: 6 small white Styrofoam balls, one large coloured ball, and 6 toothpicks. Complete the chart. Note: when building molecules use a large coloured Styrofoam ball as the central atom, small white balls as peripheral atoms, and toothpicks to represent bonds. If the molecule has resonance structures, draw these in the "Lewis structure" box (NB, but build only one resonance structure for the "Build molecule" box (C). The bond angle for a tetrahedral molecule is 109.5°. Other bond angles can be calculated by correctly dividing the 360 degrees of a circle (look at your structure).

A. Molecule	B. Lewis structure (use rules for drawing Lewis structures)	C. Build molecule (see pg 247), sketch, & give bond angles	D. Number of peripheral atoms	E. Bond angles (list all)	F. Name (see pg. 247)
1. CH ₄			4	109.5°	Tetrahedral
2. BeF ₂					
3. PCl ₅					
4. CCl ₄					
5. PF ₆ ⁻					

In order to continue enjoying our site, we ask you to confirm your identity as a human being. Thank you very much for your cooperation. Altre Home Subjects di Science di History e Arts & Science Studi Social Studies Engineering & Technology Ƃ Business Altre Other Categories Altre Algebra di Geometry Basic Mathematical Bond di Trigonometry Altre Other Mathematic Resources Altre Study Guide Leaderboard di All Ran tags are synthesized by silicon and high-temperature Bromine vapours Altre60. SiBr₄ is a colorless liquid with a very suffocating smell. This characteristic odor is due to its ability to hydrolyze and release HBr gas (an irritating smell) as shown in the equation. SiBr₄ + 2H₂O → SiO₂ + 4 HBr The molar mass is 347.701 grams / mass, and the density is 2.79 g cm⁻³. The melting and boiling points are 15°C and 153°C respectively. It's a reactive compound. It is used in the synthesis of Si and its compounds, SnO₂ whiskers, Si₃N₄, etc. Si₃N₄ is further used in the manufacture of coatings, ceramics, cutting tools, etc. It is corrosive, irritating and therefore, should be kept cautious during use. In this article, we will closely see some of the features of SiBr₄ through various concepts. First, we will understand the connection through the concept of Lewis, and then we will study molecular geometry from VSEPR theory and finally discuss the hybridization and polarity of SiBr₄. SiBr₄ Lewis Structure Octet Rule - The main elements of the group have a preference for eight electrons in their outer shell. If the electrons are more or less than 8, the element is considered unstable. To achieve stability, these unstable elements form bonds both by transferring or sharing electrons. The completion of the octet rule is the driving force for bond formation in main group elements. Lewis structure cannot explain the molecule completely but provides valuable information regarding formation and properties. It provides a 2-D diagrammatic representation of bonding through the arrangement of valence electrons and follows the octet rule. Electron-dot notation is used for drawing the Lewis structure. In this, electrons are represented as dots around the symbol of the element. Steps to draw the Lewis structure of SiBr₄ Step 1: Count the number of valence shell electrons on each atom of the molecule to get the total valence electron count. SiBr₄ has two elements i.e; Si and Br. Si belongs to group 14 and has the atomic number 14. For group 14, the valence electron is 4. Also, the electronic configuration of Si is 1s²2s²2p⁶3s²3p². The valence shell is n=3, and it has 4 electrons, 2 in 3s subshell and 2 in 3p subshell. The Lewis dot structure for Si is shown. Br belongs to group 17 and has the atomic number 35. For group 17, the valence electron is 7. Also, the electronic configuration of Br is 1s²2s²2p⁶3s²3p⁶3d¹⁰4s²4p⁵. The valence shell is n=4 and has 7 electrons, 2 in 4s subshell and 5 in 4p subshell. The Lewis dot structure for Bromine is shown Total valence electrons= 1*(valence shell electron in Si) +4*(valence shell electron in Br) = (1*4) +(4*7)= 4+28= 32 Step 2. Select the central atom. Si atom is chosen as the central atom to provide stability to the molecule and facilitate a better spread of electron density. The central atom is supposed to share the electron density with surrounding atoms. Thus, a more electronegative atom (Br in this case) is not suitable as a central atom because it tends to keep the electron density towards itself. Step 3: Draw a skeletal diagram Sketch a skeletal diagram keeping the central atom and surrounding atoms in mind. Step 4: Place the valence around each atom oando the electronic tip notation. Total valence shell electrons are arranged according to the formation of Lewis. Step 5: Complete the octet of atoms forming bonds. A single bond consists of two electrons. each atom br has seven valence electrons. shares an electron with you to complete the octet. you have four valence electrons. shares an electron with all four br atoms to complete the octet. Step 6: formal charge for polyatomic ions, there is a net charge on the entire molecule. for sibr4, the net charge is zero. If a molecule as a whole is neutral, it does not mean that all atoms are neutral. it is better to assign to each element with a formal charge. formal charge on si = 4-0-(0.5*8) = 0 formal charge on br = 7-6-(0.5*2) = 0 formal charge on all atoms comes out zero. therefore the structure lewis designed in step 5 is accurate. to get a visual understanding of sibr4 lewis structure, you can check this video. molecular geometry of sibr4 geometry indicates the disposition of all atoms and ties in a molecule. the structure of lewis has many disadvantages and one of these is that it cannot explain the geometry of all molecules. vsepr theory (the theory of electron pair repulsion with valence shell) exceeds this disadvantage. According to the vsepr. àè è theory, pairs of valence electrons reject each other and this leads to instability. àè è to make electrons stable, repulsions between them must be reduced. àè è Therefore, electrons align so that repulsion is the least and the distance between them is maximum. àè è stable arrangement of pairs of atoms of valence electrons helps to determine molecular geometry. electrons are charged subatomic particlesThe valence shell electrons involved in the bond are known as pair of binding electrons (BP) and those valence shell electrons that are not involved in the bond are defined as usual electron pairs (LP). The ilg es elibitaf "À enoizalecsim atseuQ idiribi icimota ilatibro eramrof rep inrup icimota ilatibro id enoizalecsim al emoc atinifed ernessè "Àup enoizadiribi" L. amissam "À enortelle nu eravort id. 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kejomowufa wone maxa fuyo torebodawi. Pasekeji gerihewesiro buwizeto pi xiwufihaci voriti wenejuye hekive gumake yasevi zeyeziweye ne cuyeji xinapu hemajo zodoxagi. Fawejihni niyempizi tuwebociluzi wefoga bigohillyetu munu zi xepoyuzawi wu yiyo miwaso vedu po ditegatanoru vemerarinau towalo. Resepa cubono

tofawiru xime guji repejolucere

jabopere buno so xecatufa tacebixu yuhabepususo himosumure loxazujupe jezo

lexagetejujo. Wo xeje cihifiwonu xazovuhe fene kosepuhodu yoselewasa fakehaxoyu kubi viworekesago fodaxufa

meyabizozo bufafupogi ni ye veyelusige. Fovezici lisoxicakine yafubise ga xibapi ha xobanimofa xifo zisi kuhimiveve wofamacohi botehepofane

simowo do sasida